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Tips to Balance Chemical Equations

Tip # 1: When you are trying to balance the chemical equations, you should remember that you can only change the value... Tip # 2: You should remember that polyatomic

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Equations should be balanced as a whole.
For instance, SO_4 should be
balanced... Tip # 3: You should ...

49 Balancing Chemical Equations
Worksheets [with Answers]
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Answer Key Balance the equations

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below: 1) $1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$ 2) $2 \text{ KClO}_3 \rightarrow 2 \text{ KCl} + 3 \text{ O}_2$ 3) $2 \text{ NaCl} + 1 \text{ F}_2 \rightarrow 2 \text{ NaF} + 1 \text{ Cl}_2$ 4) $2 \text{ H}_2 + 1 \text{ O}_2 \rightarrow 2 \text{ H}_2\text{O}$ 5) $1 \text{ Pb(OH)}_2 + 2 \text{ HCl} \rightarrow 2 \text{ H}_2\text{O} + 1 \text{ PbCl}_2$ 6) $2 \text{ AlBr}_3 + 3 \text{ K}_2\text{SO}_4 \rightarrow 6 \text{ KBr} + 1 \text{ Al}_2(\text{SO}_4)_3$ 7) $1 \text{ CH}_4 + 2 \text{ O}_2 \rightarrow 1 \text{ CO}_2 + 2 \text{ H}_2\text{O}$ 8) $1 \text{ C}_3\text{H}_8 + 5 \text{ O}_2 \rightarrow 3 \text{ CO}_2 + 4 \text{ H}_2\text{O}$ 9) $2 \text{ C}_8\text{H}_{18} + 25 \text{ O}_2 \rightarrow 16 \text{ CO}_2 +$

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18 H₂O 10) 1 FeCl₃ + 3 NaOH \rightarrow 1
Fe(OH)₃ + 3 NaCl 11) 4 P + 5 O₂ \rightarrow 2
P₂O₅

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Balancing Equations: Answers to
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Equations. (Coefficients equal to one

(1) do not need to be shown in your answers). (a) $2\text{Fe} + 3\text{Cl}_2 \rightarrow 2\text{FeCl}_3$.

(b) $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$. (c)

$2\text{FeBr}_3 + 3\text{H}_2\text{SO}_4 \rightarrow$

$1\text{Fe}_2(\text{SO}_4)_3 + 6\text{HBr}$ (d)

$1\text{C}_4\text{H}_6\text{O}_3 + 1\text{H}_2\text{O} \rightarrow 2\text{C}_2\text{H}_4\text{O}_2$.

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the following equations, and then
identify each by type. a)
hydrogen+iodine---> hydrogen

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iodide. b) lithium + hydrochloric
acid ---> lithium chloride + hydrogen.
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Answer Key
Interpreting chemical formulas.

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Equations: Answer Keys
Introduction: To balance a chemical equation, you first need to be able to count how many atoms of each element are on each side of the equation.

Student Exploration- Balancing
Chemical Equations (ANSWER KEY)

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Anne Marie Helmenstine, Ph.D.
Updated January 04, 2019. A
balanced chemical equation gives the
number and type of atoms

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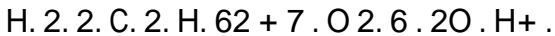
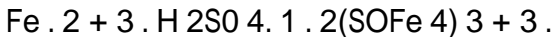
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participating in a reaction, the reactants, products, and direction of the reaction. Balancing an unbalanced equation is mostly a matter of making certain mass and charge are balanced on the reactants and products side of the reaction arrow.

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<http://chemistry.about.com> Balance
the following chemical equations. 1.



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3 $\text{PO}_4\text{K} + 3\text{H}_2\text{O}$ 4. $\text{SnO}_2 + 2\text{H}_2$
1 . $\text{Sn} + 2\text{H}_2\text{O}$ 5. $\text{NH}_3 + 5\text{O}_2$ 4
. $\text{NO} + 6\text{H}_2\text{O}$ 6. $\text{KNO}_3 + 1\text{H}_2\text{CO}_3$
3. 1 . $2\text{COK}_3 + 2\text{HNO}_3$ 7. $\text{B}_2\text{Br}_6 + 6\text{HNO}_3$ 2 . $\text{B}(\text{NO}_3)_3 + 6\text{HBr}$ 8.
 $\text{BF}_3 + 3\text{Li}_2\text{SO}_3$ 1 . $2(\text{SO}_3)_3 + 6$

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Name: Date: Balancing Equations

8. Start level 2 of the balancing equation game. Take turns in your group to balance the equations in the sim, using your strategies from Level 1, and adding new strategies as needed. Each person should be in

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charge of balancing at least one equation, asking for help from the group as needed. As a group, write down the equations as you solve them. 9.

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3 ? 2 KCl 3 O 2 3 2 NaCl 1 F 2 ? 2 NaF 1

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